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Education. Laidler received his early education at Liverpool College. He received his BA (1934) and MA (1938) degrees from Trinity College, Oxford University. His MA was in the area of chemical kinetics under Cyril Norman Hinshelwood.He completed his PhD in 1940 from Princeton University, with a thesis entitled: The Kinetics of Reactions in Condensed and

Heterogeneous Systems, under Henry Eyring.

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(i) Consider the rate law ν = k[H 2][I 2]. If we substitute units into the equation, we obtain (mol dm-3 s-1) = [k] (mol dm-3) (mol dm-3) where the notation [k] means ‘ the units of k ’. We can rearrange this expression to find the units of the rate constant, k. [k] = (mol dm-3 s-1) (mol dm-3) (mol dm-3) = mol-1 dm3 s-1

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Basic concepts of both experimental and theoretical chemical kinetics are concisely explained for those seeking a general knowledge of the subject from this well-known text, now being totally revised and updated. In addition, the book is an invaluable starting point for those embarking on research in kinetics and physical chemistry. Extensive chapter bibliographies point the way toward more detailed accounts or specialized aspects.

Historical background included in both chapter introductions and biographical sketches of important researches in chemical kinetics.

Reaction Kinetics, Volume II: Reactions in Solution deals with the kinetics of reactions in solution and discusses the basic principles and theories of kinetics, including a brief description of homogeneous gas reactions. This book is divided into two chapters. The first chapter focuses on the general principles of reactions in solution that includes reactions between ions and involving dipoles; influence of pressure on rates in solution; substituent effects; and homogeneous catalysis in solution. Chapter 2 primarily deals with general features of reactions in solution, emphasizing the relationship between the results of a kinetic investigation and actual reaction mechanism. This volume is intended for undergraduate students of chemistry who have not previously studied chemical kinetics. This book is also useful to more advanced students in other fields, such as biology and physics, who wish to have a general knowledge of the subject.

The range of courses requiring a good basic understanding of chemical kinetics is extensive, ranging from chemical engineers and pharmacists to biochemists and providing the fundamentals in chemistry. Due to the wide reaching nature of the subject readers often struggle to find a book which provides in-depth, comprehensive information without focusing on one specific subject too heavily. Here Dr Margaret Wright provides an essential introduction to the subject guiding the reader through the basics but then going on to provide a reference which professionals will continue to dip in to through their careers. Through extensive worked examples, Dr Wright, presents the theories as to why and how reactions occur, before examining the physical and chemical requirements for a reaction and the factors which can influence these. * Carefully structured, each chapter includes learning objectives, summary sections and problems. * Includes numerous applications to show relevance of kinetics and also provides plenty of worked examples integrated throughout the text.

The unusual approach of this text gives final honours and post-graduate students a clear and explanatory account of one of the " harder areas of physical chemistry. The author takes care to provide detailed verbal clarification of the concepts and their importance together with full explanations of the mathematical developments. Her explanations are an essential and vital feature of the text, which is scholarly, lucid and well-written with a combination of depth of coverage and clarity which helps students to work through on their own. A clear and explanatory account of one of the more difficult areas of physical chemistry Provides

detailed verbal clarification of the concepts and their importance together with full explanations of the mathematical developments Discusses energy transfer, molecular beam studies of reactive scattering and historical developments and modern kinetics, among other topics

Unimolecular reactions are in principle the simplest chemical reactions, because they only involve one molecule. The basic mechanism, in which the competition between the chemical reaction step and a collisional deactivation leads to a pressure-dependent coefficient, has been understood for a long time. However, this is a rapidly developing field, and many new and important discoveries have been made in the past decade. This First Part Part of Two CCK Volumes dealing with Unimolecular Rections, deals with the Reaction Step. The first chapter is an introduction to the whole project, aiming to cover the material necessary to understand the content of the detailed chapters, as well as the history of the development of the area. Chapter 2 is a review of the modern view of the statistical theories, as embodied in the various forms of RRKM theory. Chapter 3 deals with the fully quantum mechanical view of reactive states as resonances. . Presents considerable advances in the field made during the last decade. . Treats both the statistical as well as the fully quantum mechanical view.